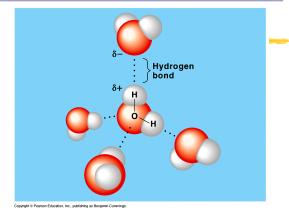
CHAPTER 3 WATER & ENVIRONMENT



Figure 3.1 Hydrogen bonds between water molecules



BONDING OF WATER

- OXYGEN HAS SLIGHT CHARGE
- HYDROGEN HAS + CHARGE
- POLAR MOLECULE FORMED
- BRING ABOUT SPECIAL PROPERTIES
- WTER MOLECULES CAN ATTACH TO 4 OTHER WATER MOLECULES

Figure 3.5 The structure of ice (Layer 1)

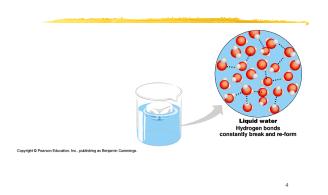




Figure 3.5 The structure of ice (Layer 2)

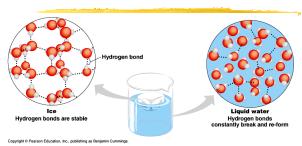
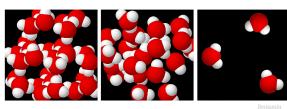


Figure 3.5x1 Ice, water, and steam



Benjamin Cumming

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PROPERTIES OF WATER

- CHOESION OF MOLECULES
- ADHESION OF MOLECULES
- SURFACE TENSION
- CAPILLARITY
- ATTRACTIONDISSOCAIATION
- UNLIKE MOLECULES

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- LIKE MOLECULES
- ADHESION AND
- COHESION

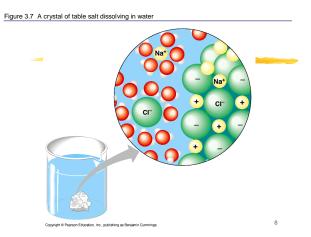


Figure 3.2 Water transport in plants



WATER/TEMPERATURE

- HIGH SPECIFIC HEAT EVAPORATIVE
- STABALIZES TEMPERATURE
- RESISTS CHANGE OF
- COOLING MOLECULES BECOME UNHOOKED
- LOOSES OR ABSORBS HEAT OF HEAT VAPORIZ
 - VAPORIZATION STABLIZES FROM OVERHEATING

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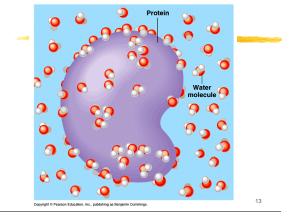
WATER AND ICE

- ICE FLOATS
- WATER COOLS AND EXPANDS
- WATER IS LESS DENSE AS A SOLID
- BODIES OF WATER THEN TEND NOT TO FREEZE SOLID
- HELPING OUR BODIES?

WATER /SOLVENT/LIFE

- WATER IS UNIVERSAL SOLVENT
- SOLUTES DISSOLVE IN WATER
- IONIC SOLUTIONS DISSOLVE IN WATER BECAUSE OF ELECTRICAL AFFINITY OF SALT AND WATER
- HYDROPHILIC/PHOBIC SUBSTANCES

Figure 3.8 A water-soluble protein

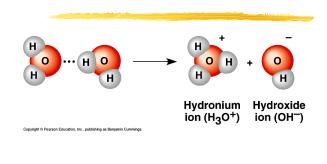




WATER AND pH

- pH Scale is from 0-14
- Measure H+ ion number
- Iow pH, high # of H ions
- high pH, low # of H ions
- Living systems operate best in pH of 6.8-7.2, blood is pH7.5
- Buffers balance pH. Absorb ions

Unnumbered Figure (page 47) Chemical reaction: hydrogen bond shift



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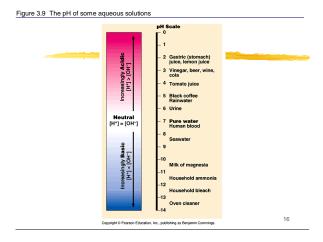




Figure 3.10 The effects of acid precipitation on a forest



Figure 3.10x1 Pulp mill



Figure 3.10x2 Acid rain damage to statuary, 1908 & 1968



